

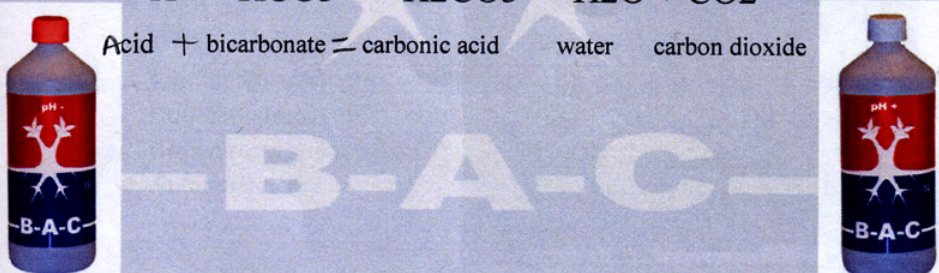
pH raises because bicarbonates and Carbonic acid
 cause it to be unstable
 lower pH to 5.3 to release bicarbonate, then add pH up

pH

Influences: light temp, air

- pH = amount of acid (H⁺)
- pH tap water
 - $H^+ + HCO_3^- \rightarrow H_2CO_3 \rightarrow H_2O + CO_2$

Acid + bicarbonate = carbonic acid water carbon dioxide




-B-A-C-

Coca Cola
 fizzy evaporates,
 then
 Carbonic acid
 leaves coke

if bicarbonate is at 5.3 it will collapse

pH

- Absorption at the root
 - Positive ions: Pota. + ions --> pH lowering
 - Negative ions: Nitr. en Phos - ions --> pH increase
- Water pH (tank)
- Plant formula / pH
- Nutrients pH (substrate root environment)
- Balance principle



-B-A-

K = pH down
 N + P = pH up

Potassium bicarbonate
 Potassium hydroxide